


UNIVERSITY OF MYSORE
Estd. 1916

No.AC2(S)/151/2020-21

VishwavidyanilayaKaryasoudha
Crawford Hall, Mysuru- 570 005

Dated: 26-10-2021

Notification

Sub:- Syllabus and Examination Pattern of Chemistry (UG) with effective from the Academic year 2021-22 as per NEP-2020.

- Ref:-**
1. Decision of Board of Studies in Chemistry (UG) meeting held on 01-10-2021.
 2. Decision of the Faculty of Science & Technology Meeting held on 16-10-2021.
 3. Decision of the Academic Council meeting held on 22-10-2021.

The Board of studies in Chemistry (UG) which met on 01-10-2021 has recommended & approved the syllabus and pattern of Examination of Chemistry Programme with effective from the Academic year 2021-22 as per NEP -2020.

The Faculty of Science & Technology and Academic Council at their meetings held on 16-10-2021 and 22-10-2021 respectively have also approved the above said proposal and it is hereby notified.

The syllabus and Examination pattern is annexed herewith and the contents may be downloaded from the University Website i.e., www.uni-mysore.ac.in.


Registrar
Registrar
University of Mysore
Mysore

To:-

1. All the Principal of affiliated Colleges of University of Mysore, Mysore. Those who are running B.Sc Courses.
2. The Registrar (Evaluation), University of Mysore, Mysuru.
3. The Chairman, BOS/DOS, in Chemistry, Manasagangothri, Mysore.
4. The Dean, Faculty of Science & Technology, DoS in Psychology, MGM.
5. The Director, Distance Education Programme, Moulya Bhavan, Manasagangothri, Mysuru.
6. The Director, PMEB, Manasagangothri, Mysore.
7. Director, College Development Council , Manasagangothri, Mysore.
8. The Deputy Registrar/Assistant Registrar/Superintendent, Administrative Branch and Examination Branch, University of Mysore, Mysuru.
9. The PA to Vice-Chancellor/ Registrar/ Registrar (Evaluation), University of Mysore, Mysuru.

UNIVERSITY OF MYSORE



B.Sc. (Honors) Degree Programme in Chemistry

NATIONAL EDUCATION POLICY (NEP) – 2020

**CHOICE BASED CREDIT SYSTEM (CBCS) WITH
MULTIPLE ENTRY AND EXIT OPTIONS**

2021-22

Program Outcomes:

By the end of the program the students will be able to:

(Refer to literature on outcome-based education (OBE) for details on Program Outcomes)

1. **PO. 1:** To create enthusiasm among students for chemistry and its application in various fields of life.
2. **PO. 2:** To provide students with broad and balanced knowledge and understanding of key concepts in chemistry
3. **PO. 3:** To develop in students a range of practical skills so that they can understand and assess risks and work safely measures to be followed in the laboratory.
4. **PO. 4:** To develop in students the ability to apply standard methodology to the solution of problems in chemistry
5. **PO. 5:** To provide students with knowledge and skill towards employment or higher education in Analytical chemistry or multi-disciplinary areas involving chemistry.
6. **PO.6:** To provide students with the ability to plan and carry out experiments independently and assess the significance of outcomes and to cater to the demands of chemical Industries of well-trained graduates
7. **PO. 7:** To develop in students the ability to adapt and apply methodology to the solution of unfamiliar types of problems.
8. **PO. 8:** To instill critical awareness of advances at the forefront of chemical sciences, to prepare students effectively for professional employment or research degrees in chemical sciences and to develop an independent and responsible work ethics.

Details of the B.Sc (Honors) degree programme in Chemistry

Semester	Course Type								DSE				OE						
	DSC																		
	THEORY	L	T	P	PRACTICALS	L	T	P		L	T	P	L	T	p				
I	A1	4	0	0	P-1	0	0	2					3	0	0				
II	A2	4	0	0	P-2	0	0	2					3	0	0				
III	A3	4	0	0	P-3	0	0	2					3	0	0				
IV	A4	4	0	0	P-4	0	0	2					3	0	0				
V	A5	3	0	0	P-5	0	0	2	A1	3	0	0							
	A6	3	0	2	P-6	0	0	2											
VI	A7	3	0	0	P-7	0	0	2	A2	3	0	0							
	A8	3	0	0	P-8	0	0	2											
VII	A9	3	0	0	P9	0	0	2	A3	3	0	0							
	A10	3	0	0	P10	0	0	2								RESEARCH METHADODOLOGY	3	0	0
	A11	4	0	0															
VIII	A12	4	0	0					A4	3	0	0							
	A13	4	0	0												PROJECT	0	0	6
	A14	3	0	0															
TOTAL CREDITS		69							21										

DSC: DISCIPLINE SPECIFIC COURSE

OE: OPEN ELECTIVE.

DSE: DISCIPLINE SPECIFIC ELECTIVE

L : T : P = Lecture : Tutorial : Practical

GENERAL REQUIREMENTS AND OTHER INFORMATIONS.

Scheme of Instructions

1. **Title and Commencement:** As per the university guidelines (12 Ref.letterUA2/379/2016-17).
2. Undergraduate programme offered with multiple entry and exit options
Faculty of Science–
Certificate – 2 semesters
Diploma– 4semesters
Bachelor of Science (B.Sc. 6 Semesters)
Bachelor of Science. Honors (B.Sc.Hons, 8 Semesters)

3. Semester and ProgrammeStructure:

The credit pattern for the course is L:P

3. **Subject Combinations:** As per the university guidelines (Ref. letterUA2/379/2016-17).

4. Eligibility for Admission.

For B.Sc program only those students who have completed PUC with chemistry or its equivalent examination with science subjects are eligible.

5. **Medium of Instruction:** The medium of instruction shall beEnglish/Kannada.
6. **Scheme of the Program:** As per the university guidelines (Ref.letterUA2/379/2016-17).
7. **Course Registration:** As per the university guidelines (9.1 to 9.6 Ref.letterUA2/379/2016-17).
8. **Attendance:** As per the university guidelines (10.1 and 10.2 Ref.letterUA2/379/2016-17)
9. **Valuation:**Aspertheuniversityguidelines(Ref.letterAC2(S)/151/2021-22,dated18/08/2021
10. Ifthestudenthaspassedinthepracticalexambysecuringprescribedmarksneednotreappear for the practical exam if he/she has failed in the theoryexam.

11. PassingCriteria

A student is considered to have passed the course, only on securing a minimum of 40% from internal assessment and end examination marks put together.

A student can take end exam irrespective of the marks scored in internal assessment of a particular course

In case a student secures less than 30% in end exam or absent for end examination, the student is said to have not completed the course. The student shall complete the course by reappearing only the end examination conducted by the university.

Makeup examination: As per the university guidelines (16. Ref. letterUA2/379/2016-17).

Percentage and Grading: As per the university guidelines (17 Ref. letter UA2/379/2016-17). **18 to 22.** As per the university guidelines (Ref. letterUA2/379/2016-17)

**Scheme of Examination for DSC 1 and 2
(I and II Semester)**

Credits L : P		Theory	Practical	Maximum marks
4 : 0	Internal assessment	40 C1 = 10 + 10 = 20 C2 = 10 + 10 = 20 (test and assignment)	25 C1 = 10 C2 = 10 + Record 5	
0 : 2	Summative Assessment	60 (C3)	25	
Duration of the end examination		2 hours	4 hours	
		100	50	150

Examination and Evaluation

**Question paper pattern for DSC 1 and 2
(I and II Semester)**

Duration: 2 hours		Max. Marks: 60
The question paper contains 3 parts		
Part-A	Answer any 6 out of 8 questions (two questions from each unit)	6 X 2 = 12
Part-B (Analytical Chemistry)	Answer any 2 out of 3 questions	2 x 6 = 12
Part-C (Inorganic Chemistry)	Answer any 2 out of 3 questions	2 x 6 = 12
Part-D (Organic Chemistry)	Answer any 2 out of 3 questions	2 x 6 = 12
Part-E (Physical Chemistry)	Answer any 2 out of 3 questions	2 x 6 = 12
Pattern: (3 + 3) / (4 + 2)/(2+2+2)		

Scheme of Examination for Open elective

Credits L : P	Theory	Maximum marks
3: 0	Internal assessment	40 C1 = 10 + 10 = 20 C2 = 10 + 10 = 20 (test and assignment)
	Summative assessment	60 (C3)
Duration of the end examination		2 hours
		100

Question paper pattern for Open elective (I to IV Semester)

Duration: 2 hours		Max. Marks: 60
The question paper contains 2 parts		
Part-A	Answer any 6 out of 8 questions	6 X 2 = 12
Part-B	Answer any 8 out of 10 questions	8 x 6 = 48
Pattern: Pattern: (3 + 3) / (4 + 2)/(2+2+2)		

SCHEME OF VALUATION IN CHEMISTRY FOR THE
PRACTICALEXAMINATIONS2021-22

I Semester

Inorganic and organic chemistry practicals P-1

Chemistry DSC-P-1(Practical-I)

Max marks: 50

Practicaltest : 20 marks

Record : 05marks

FinalPracticalexamination : 25marks

Practical duration: 4Hours

Note: Duly Certified practical record shall be submitted at the practical examination (No evaluation of record)

Final Practical Examination:

Part A – Volumetric estimation (Inorganic)	13 marks
Part B – Organic preparation	12 marks

Part A- volumetric estimation

Scheme for volumetric estimation (Experiments 1,5,6) – 13 Marks

- a) Preparation of Standard solution and Calculation of Normality: (1+1 marks)=02marks
b) Titre values of Standardization and estimation (3 + 5) = 8marks

Experimental Values:

Experimental values	Marks for standardization	Marks for estimation
±0.2 cm ³	3	5
±0.3 cm ³	2	4
Any other value	1	3

- c) Calculation:(Normality of link solution, given solution and Wt/dm³ or 500cm³ or 250cm³ (1+1+1marks)3marks

Scheme for volumetric estimation (Experiments 2,3,4) – 13 Marks

Discrepancy	First titration	Second titration
$\pm 0.2 \text{ cm}^3$	4	5
$\pm 0.3 \text{ cm}^3$	3	4
Any other value	2	3

Calculation: 2 +2

Part – B – Organic preparations. – 12 Marks

Scheme for organic preparation

Preparation – 06Marks

Equation – 02 Marks

Yield – 01 Marks

Recrystallization – 02 Marks

Melting point – 01 Marks

II Semester
Analytical and physical Chemistry practicals P-2

Chemistry DSC-P-2(Practical-II)

Max. marks:50

Practical test : 20 marks

Record : 05 marks

Final Practical examination : 25 marks

Practical duration: 4 Hours

Note: Duly Certified practical record shall be submitted at the practical examination
(No evaluation of record)

Final Practical Examination:

Part A – Volumetric Estimation (Analytical)	13 Marks
Part B – Physical Chemistry Experiments	12 Marks

Part A – Analytical Chemistry experiments

Volumetric estimation

13 Marks

Scheme of valuation (Experiments 1,3,4,5)

- a) Preparation of Standard solution and Calculation of Normality: (1+1 marks)=02 marks
- b) Titre values of Standardization and estimation (3 + 5) = 8 marks

Experimental Values:

Experimental values	Marks for standardization	Marks for estimation
±0.2 cm ³	3	5
±0.3 cm ³	2	4
Any other value	1	3

- c) Calculation: (Normality of link solution, given solution and wt/dm³ or 500cm³ or 250cm³ (1+1+1 marks) 3 marks

Scheme (Experiment 2) – 13 Marks

- a) Preparation of Standard $K_2Cr_2O_7$ solution: = 02marks
- b) calculation of Normality – 2marks
- c) Titre values of estimation = 6 marks

Experimental Values:

Experimental values	Marks for estimation
$\pm 0.2 \text{ cm}^3$	5
$\pm 0.3 \text{ cm}^3$	4
Any other value	3

- d) Calculation: (Normality of given solution and wt/dm³ or 500cm³ or 250cm³
(2+1marks)3marks

Scheme (Experiment 6) – 13 Marks

- a) Preparation of two solution(2+2) : = 04marks
- b) estimation 1 and 2

Experimental Values:

Experimental values	Estimation 1	Estimation 2
$\pm 0.2 \text{ cm}^3$	4	4
$\pm 0.3 \text{ cm}^3$	3	3
Any other value	2	2

- c) Calculation: (Normality of given solution and wt/dm³ or 500cm³ or 250cm³ = 1marks

Part B – Physical chemistry experiments

Distribution of marks: 12

Sl. No.	Experiment	Discrepancy between and Reportedvalue	Marks	Experiment.	Discrepancy between and Reportedvalue	Marks
1	Density	± 0.2	02	Viscosity	± 0.5	06
		± 0.3	01		± 1.0	05
		Any other Value	01		Any other Value	04
		Calculation	01		Calculation	01
		SI unit	01		SI unit	01
2.	Density	± 0.2	02	Surface tension	± 0.5	06
		± 0.3	01		± 1.0	05
		Any other Value	01		Any other Value	04
		Calculation	01		Calculation	01
		SI unit	01		SI unit	01
3.	Percentage composition of given mixture using Abbe'S Refractometer.	Preparation of Mixture $\pm 5\%$. $\pm 10\%$. Any other value Graph	03 07 06 05 02			
4	Determination of partition coefficient	(i) (ii) (iii)	4 4 4			

5	Kinetics of hydrogen peroxide	5 constant values 4 constant values Any other value Calculation Graph log V2/V1	07 06 05 02 02			
6	Percentage of NaCl Unknown(one)	preparation of mixture ± 0.2 % ± 0.3 % Any other value Graph	04 06 05 04 02			

I SEMESTER

DSC-1: Chemistry-1

CLASS DURATION – THEORY: 04 HOURS/WEEK

Theory and Practicals: Total Credits-06 (Theory-04, Practicals-02)

UNIT –I – Analytical chemistry

Language of analytical chemistry: Definitions of analysis, determination, measurement, techniques and methods. Classification of analytical techniques. Choice of an analytical method - accuracy, precision, sensitivity, selectivity, method validation. Figures of merit of analytical methods and limit of detection (LOD), Limit of quantification (LOQ), linear dynamic range (working range).

Errors and treatment of analytical data: Limitations of analytical methods – Errors: Determinate and indeterminate errors, absolute error, relative error, minimization of errors. Statistical treatment of finite samples -mean, median, range, standard deviation and variance. External standard calibration - regression equation (least squares method), correlation coefficient (R²).

Numerical problems

Basic laboratory practices, calibration of glassware (pipette, burette and volumetric flask), Sampling (solids and liquids), weighing, drying, dissolving, Acid treatment, Rules of work in analytical laboratory, General rule for performing quantitative determinations (volumetric and gravimetric), Safety in Chemical laboratory, Rules of fire prevention and accidents, First aid. Precautions to be taken while handling toxic chemicals, concentrated/fuming acids and organic solvents. **[14 Hours]**

UNIT-II: Inorganic Chemistry

Bohr's theory, its limitations and atomic spectrum of hydrogen atom. Wave mechanics: de Broglie equation, Heisenberg's Uncertainty Principle and its significance, Schrödinger's wave equation, significance of ψ and ψ^2 . Quantum numbers and their significance.

Normalized and orthogonal wave functions. Sign of wave functions. Radial and angular wave functions for hydrogen atom. Radial and angular distribution curves. Shapes of s, p, d and f orbitals. Contour boundary and probability diagrams.

Pauli's Exclusion Principle, Hund's rule of maximum multiplicity, Aufbau's principle and its limitations-Electronic configurations of the elements (Z=1-30), effective nuclear charge, shielding/screening effect, Slater's rules. Variation of effective nuclear charge in Periodic Table. **[14 Hours]**

UNIT-III: Organic Chemistry

Classification and nomenclature of organic compounds, Hybridization, Shapes of organic molecules, Influence of hybridization on bond properties.

Nature of bonding in Organic molecules

Formation of Covalent bond, Types of chemical bonding, localized and delocalized, conjugation and cross conjugation, concept of resonance, electronic displacements: Inductive effect, Electromeric effect, Resonance and Hyper conjugation, cross conjugation explanation with examples. Concept of resonance, aromaticity, Huckel rule, anti-aromaticity explanation with examples. Strengths of Organic acid and bases: Comparative study with emphasis on factors effecting pK values. Relative strength of aliphatic and aromatic carboxylic acids-Acetic acid and chloroacetic acid, acetic acid and propionic acid, acetic acid and Benzoic acid. Steric effect- Relative stability of trans and cis-2-butene.

Mechanisms of Organic Reactions

Notations used to represent electron movements and directions of reactions- curly arrows, formal charges. Types of bonds breaking- homolytic and heterolytic. Types of reagents-Electrophiles, nucleophiles, nucleophilicity and basicity. Types of organic reactions- substitution, addition, elimination, rearrangement and pericyclic reactions, explanation with examples.

Chemistry of Aliphatic hydrocarbons, Carbon-Carbon Sigma bonds

Chemistry of alkanes: Formation of alkanes, Wurtz reaction, Wurtz-Fittig reaction, Free radical substitution, Halogenation- relative reactivity and selectivity

Carbon-carbon pi bonds

Formation of alkenes and alkynes by elimination reaction. Mechanism of E1, E2, E1cb reaction. Saytzeff and Hofmann eliminations. Addition of HBr to propene, Free radical addition of HBr to propene. Addition of halogens to alkenes-carbocation and halonium ion mechanism. Stereospecificity of halogen addition. Ozonolysis mechanism - ozonolysis of propene. Addition of hydrogen halides to alkenes, mechanism, regioselectivity and relative rates of addition. Hydrogenation, hydration, hydroxylation and epoxidation of alkenes, explanation with examples, 1,2 and 1,4- addition reactions in conjugated dienes. Diels-Alder reaction, Allylic and benzylic bromination and mechanism in propene, 1-butene, 1-toluene and ethylbenzene.

[14 hours]

UNIT-IV: Physical Chemistry

Gaseous State

Elementary aspects of kinetic theory of gases, Ideal and real gases. Boyle temperature (derivation not required), Molecular velocity, collision frequency, collision diameter, Collision cross section, collision number and mean free path and coefficient of viscosity, calculation of σ and η , variation of viscosity with temperature and pressure.

Maxwell's Boltzmann distribution law of molecular velocities (Most probable, average and root mean square velocities). Relation between RMS, average and most probable velocity and average kinetic energies. (Mathematical derivation not required), law of equipartition of energy.

Behaviour of real gases: Deviation from ideal gas behaviour. Compressibility factor (Z) and its variation with pressure for different gases. Causes of deviation from ideal behaviour, vanderWaals equation of state (No derivation) and application in explaining real gas behaviour. Critical phenomena - Andrews isotherms of CO₂, critical constants and their calculation from van der Waals equation, Continuity of states, Law of corresponding states. Numerical problems.

Liquid State

Surface Tension: Definition and its determination using stalagmometer, effect of temperature and solute on surface tension

Viscosity: Definition, Coefficient of viscosity. Determination of viscosity of a liquid using Oswald viscometer. Effect of temperature, size, weight, shape of molecules and intermolecular forces.

Refraction: Specific and molar refraction- definition and advantages. Determination of refractive index by Abbes Refractometer. Additive and constitutive properties.

Parachor: Definition, Atomic and structure parachor, Elucidation of structure of benzene and benzoquinone. Viscosity and molecular structure. Molar refraction and chemical constitution. Numerical Problems.

[14 Hours]

I Semester Practicals

CHEMISTRY-DSC 1 LAB: 04 HOURS/WEEK

Content of Practical Course 1: List of Experiments

PART-A Inorganic Chemistry

1. Preparation of standard sodium carbonate solution and standardization of hydrochloric acid solution (methyl orange indicator). Estimation of sodium hydroxide present in the solution using phenolphthalein indicator.
2. Determination of carbonate and hydroxide present in a mixture.
3. Determination of oxalic acid and sodium oxalate in a given mixture using standard $\text{KMnO}_4/\text{NaOH}$ solution
4. Estimation of ferrous and ferric iron in a given mixture using standard potassium dichromate solution
5. Preparation of standard oxalic acid solution and standardization of potassium permanganate solution. Estimation of hydrogen peroxide present in the solution.
6. Preparation of standard oxalic acid solution and standardization of potassium permanganate solution. Estimation of ferrous ammonium sulphate present in the solution.

PART-B Organic Chemistry

1. Preparation of acetanilide from aniline using $\text{Zn}/\text{acetic acid}$ (Green method).
2. Synthesis of p-nitro acetanilide from acetanilide using nitrating mixture.
3. Bromination of acetanilide
4. Hydrolysis of methyl m-nitrobenzoate to m-nitrobenzoic acid (Conventional method)
5. Synthesis of diazoaminobenzene from aniline (conventional method).
6. Preparation of dibenzalacetone (Green method).

LEARNING OUTCOMES / COURSE OUTCOMES

Chemistry as Discipline Specific Course (DSC)

B.Sc. Semester – I

CHEMISTRY:1

- The concepts of chemical analysis, accuracy, precision and statistical data treatment
- Prepare the solutions after calculating the required quantity of salts in preparing the reagents/solutions and dilution of stock solution.
- Describe the dual nature of radiation and matter; dual behavior of matter and radiation, de Broglie's equations, Heisenberg uncertainty principle and their related problems.
- Quantum mechanics. Derivation of Schrodinger's wave equation. Radial and angular Orbital shapes of s, p, d and f atomic orbitals, nodal planes. Electronic configurations of the atoms.
- Pauli's exclusion principle, Hund's rule, Aufbau's principle and its limitation.
- The concepts of Organic reactions and techniques of writing the movement of electrons, bond breaking, bond forming
- The Concept of aromaticity, resonance, hyper conjugation, etc.
- Explain bond properties, electron displacement effects (inductive effect, electrometric effect, resonance effect and Hyper conjugation effect). steric effect and their applications in explaining acidic strength of carboxylic acids, basicity of amines.
- Understand basic concept of organic reaction mechanism, types of organic reactions.
- Understand the preparation and reactions of alkanes.
- Understand the stability and conformational analysis of cycloalkanes.
- Understand the concept of resonance, aromaticity and antiaromaticity.
- Describe relative strength of aliphatic and aromatic carboxylic acids.
- Explain the existence of different states of matter in terms of balance between intermolecular forces and thermal energy of the particles. Explain the laws governing behavior of ideal gases and real gases. Understand cooling effect of gas on adiabatic expansion
- Understand the conditions required for liquefaction of gases. Realize that there is continuity in gaseous and liquid state.
- Understand the properties of liquids in terms of intermolecular attractions.
- Understand the existence of different states of matter in terms of balance

between intermolecular forces and thermal energy of the particles. Explain the laws governing behavior of ideal gases and real gases. Understand cooling effect of gas on adiabatic expansion

- Understand the conditions required for liquefaction of gases. Realize that there is continuity in gaseous and liquid state.
- Understand the properties of liquids in terms of intermolecular attractions.

CHEMISTRY LAB (volumetric (inorganic) and Organic preparations):P-1

After studying this course and performing the experiments set in it student will be able to:

1. Basic concepts involved in titrimetric analysis, primary standard substances, preparation of standard solutions.
2. Explain the principles of acid-base, redox and iodometric titrations.
3. Work out the stoichiometric relations based on the reactions involved in the titrimetric analysis.
4. Understand the preparation of organic compounds involving addition, substitution, hydrolysis, diazotization and condensation reactions.

II SEMESTER

DSC-2: Chemistry-2

CLASS DURATION – THEORY: 04 HOURS/WEEK

Theory and Practicals: Total Credits-06 (Theory-04, Practicals-02)

UNIT-I: Analytical Chemistry

Titrimetric analysis: Basic principle of titrimetric analysis. Classification, Preparation and dilution of reagents/solutions. Normality, Molarity and Mole fraction. Use of $N_1V_1 = N_2V_2$ formula, Preparation of ppm level solutions from source materials (salts), conversion factors.

Acid-basetitrimetry: Titration curves for strong acid vs strong base, weak acid vs strong base and weak base vs strong acid titrations. Titration curves, Quantitative applications – selecting and standardizing a titrant, inorganic analysis - alkalinity, acidity.

Complexometric titrimetry: Indicators for EDTA titrations - theory of metal ion indicators, titration methods employing EDTA - direct, back, displacement and indirect determinations, Application-determination of hardness of water.

Redox titrimetry: Balancing redox equations, calculation of the equilibrium constant of redox reactions, titration curves, Theory of redox indicators, calculation of standard potentials using Nernst equation. Applications.

Precipitation titrimetry: Titration curves, titrants and standards, indicators for precipitation titrations involving silver nitrate- Volhard's and Mohr's methods and their differences.

Gravimetric Analysis: Requisites of precipitation, mechanism of precipitation, Factors influencing precipitation, Co-precipitation, post-precipitation, Advantages of organic reagents over inorganic reagents, reagents used in gravimetry (8-hydroxy quinoline (oxine) and dimethyl glyoxime (DMG).

Numerical problems on all the above aspects.

[14 hours]

Unit – II Inorganic chemistry

s, p, d and f-block elements, the long form of periodic table. Detailed discussion of the following properties of the elements, with reference to s and p-block elements:

(a) Atomic radii (van der Waals)

(b) Ionic and crystal radii.

(c) Covalent radii

(d) Ionization enthalpy, successive ionization enthalpies and factors affecting ionization energy. Applications of ionization enthalpy.

(e) Electron gain enthalpy, trends of electron gain enthalpy.

(f) Electronegativity, Pauling's/ Mulliken's/ Allred Rachow's/ and Mulliken-Jaffé's electronegativity scales. Variation of electronegativity with bond order, partial charge, hybridization, group electronegativity.

Trends in the chemistry of the compounds of groups 13 to 17 (hydrides, carbides, oxides and halides) are to be discussed.

[14 hours]

Unit – III Organic chemistry

Nucleophilic substitution at saturated carbon. Mechanism of SN1 and SN2 reactions with suitable examples. Energy profile diagrams, Stereochemistry and factors effecting SN1 and SN2 reactions.

Aromatic Electrophilic substitution reactions, Mechanisms, σ and π complexes, Halogenation, Nitration,

Sulphonation, Friedel Crafts alkylation and acylation with their mechanism. Activating and deactivating groups. Orientation influence, Ortho-para ratio.

Aromatic nucleophilic substitution reaction: S_NAr and Benzyne mechanism with suitable examples

[14 hours]

Unit – IV Physical Chemistry

Liquid Crystals

Explanation, classification with examples- Smetic, nematic, cholesteric, discs shaped and polymeric. Structures of nematic and cholesteric phases-molecular arrangements in nematic and cholesteric liquid crystals. Applications of liquid crystals in LCDs and thermal sensing.

Solids

Forms of solids: Unit cell and space lattice, anisotropy of crystals, size and shape of crystals,

Laws of Crystallography: Law of constancy of interfacial angles, Law of rational indices, Law of symmetry (Symmetry elements), Crystal systems, Bravais lattice types and identification of lattice planes.

Miller indices and its calculation, X-Ray diffraction by crystals: Bragg's law and derivation of Bragg's equation, Single crystal and powder diffraction methods. Defects in crystals, glasses and liquid crystals. Numerical problems.

Distribution Law

Nernst Distribution Law - Statement and its derivation. Distribution constant, factors affecting distribution constant, validity of Distribution Law, Modification of distribution law when molecules undergo a) Association b) Dissociation. Application of Distribution Law in Solvent extraction. Derivation for simple and multiple extraction. Principles of distribution law in Parkes Process of desilverisation of lead. Numerical Problems.

[14 hours]

II Semester Practicals

CHEMISTRY-DSC 2 LAB: 04HOURS/WEEK

Content of Practical Course 2: List of Experiments

PART-A Analytical Chemistry

1. Preparation of standard sodium carbonate solution; standardization of given HCl solution and estimation of alkali present in soap/detergent
2. Preparation of standard K₂Cr₂O₇ solution and estimation of Iron (II) in the given solution.
3. Preparation of standard oxalic acid solution standardization of given KMnO₄ solution and estimation of given oxalic acid solution.
4. Preparation of EDTA solution and estimation of hardness (CaCO₃) of two different samples.
5. Preparation of standard Na₂CO₃ solution , standardization of given HCl solution and estimation of alkali present in given antacid.
6. Determination of chlorine in two different samples of bleaching powder by iodometry (standard sodium thiosulphate solution to be supplied)

PART-B Physical Chemistry

1. Determination of density using specific gravity bottle and viscosity of liquids using Ostwald's viscometer (Ethyl acetate, Toluene, Chloroform, Chlorobenzene or any other non-hazardous liquids)
2. Determination of the density using specific gravity bottle and surface tension of liquids using Stalagmometer (Ethyl acetate, Toluene, Chlorobenzene, any other non-hazardous liquids)
3. Determination of the composition of liquid mixture by refractometry. (Toluene & Alcohol, Water & Sucrose)
4. Determination of partition/distribution coefficient - i) Acetic acid in water and cyclohexane. ii) Acetic acid in Water and Butanol. iii) Benzoic acid in water and toluene.
5. Determination of rate constant of decomposition of H₂O₂ catalyzed by FeCl₃
6. Determination of percentage composition of NaCl solution by determining miscibility temperature of phenol-water system.

LEARNING OUTCOMES / COURSE OUTCOMES

Chemistry as Discipline Specific Course (DSC)

B.Sc. Semester II

CHEMISTRY:2

- Understand principles of titrimetric analysis.
- Understand principles of different type's titrations. Titration curves for all types of acids – base titrations.
- Gain knowledge about balancing redox equations, titration curves, theory of redox indicators and applications.
- Understand titration curves, indicators for precipitation titrations involving silver nitrate- Volhard's and Mohr's methods and their differences.
- Indicators for EDTA titrations - theory of metal ion indicators. Determination of hardness of water.
- Understand periodic table, classification and properties of s p d and f block elements
- Understand different scales for the measurement of electro-negativity and factors affecting it.
- Understand the chemistry of the hydrides, carbides, oxides and halides of group 13 to 17
- Understand nucleophilic substitution at saturated carbon, energy profile diagram stereochemistry and factors affecting S_N¹ and S_N² reactions.
- Aromatic electrophilic substitution reactions like nitration sulphonation Friedel-Crafts reactions

- Understand liquid crystals, classification with examples
- Understand the different forms of solids, laws of crystallography , miller indices and its calculation, X-ray diffraction studies. Bragg's law and its equation
- Defects in solids , properties of glasses and concept of liquid crystals

CHEMISTRY LAB - Volumetric (analytical) and physical practicals):P-2

After studying this course and performing the experiments set in it student will be able to:

1. Basic concepts involved in titrimetric analysis, primary standard substances, preparation of standard solutions.
2. Explain the principles of acid-base, redox and iodometric titrations.
3. Describe the significance of inorganic quantitative analysis.
4. Determine density followed by the determination of viscosity and surface tension of different liquid samples.
5. Determination of partition coefficient of different liquid mixtures
6. Determination of rate constant in the decomposition reaction of hydrogen peroxide

Reference Books for Discipline Specific Course

Analytical Chemistry

1. Jeffery, G.H., Bassett, J., Mendham, J. & Denney, R.C.
2. Vogel's Textbook of Quantitative Chemical Analysis, John Wiley & Sons, 1989.
3. Willard, H. H., Merritt, L.L., Dean, J. & Settle, F.A. Instrumental Methods of Analysis, 7th Ed. Wadsworth Publishing Company Ltd., Belmont, California, USA, 1988.
4. Christian, G.D; Analytical Chemistry, VI Ed. John Wiley & Sons, New York, 2004.
5. Harris, D. C. Exploring Chemical Analysis, Ed. New York, W.H. Freeman, 2001.
6. Skoog, D. A. Holler F.J. & Nieman, T.A. Principles of Instrumental Analysis, Cengage Learning India Ed, 2017.
7. Ditts, R.V. Analytical Chemistry; Methods of Separation, van Nostrand, 1974.

Inorganic Chemistry

1. Lee, J.D. Concise Inorganic Chemistry ELBS, 1991.
2. Cotton, F.A., Wilkinson, G. & Gaus, P.L. Basic Inorganic Chemistry, 3rd Ed., Wiley.
3. Douglas, B.E., McDaniel, D.H. & Alexander, J. J. Concepts and Models in Inorganic Chemistry, John Wiley & Sons.
4. Huheey, J. E., Keiter, E.A., Keiter, R.L. & Medhi, O. K. Inorganic Chemistry: Principles of Structure and Reactivity, Pearson Education India, 2006.
5. Shriver, D.F. & Atkins, P.W. Inorganic Chemistry, Oxford University Press.
6. Wulfsberg, G. Inorganic Chemistry, Viva Books Pvt.Ltd.
7. Rodgers, G. E. Inorganic & Solid State Chemistry, Cengage Learning India Ltd., 2008.
8. Mark Weller and Fraser Armstrong, 5th Edition, Oxford University Press (2011-2012)
Adam, D.M. Inorganic Solids: An introduction to concepts in solid-state structural chemistry. John Wiley & Sons, 1974.
9. G.L. Miessler & Donald A. Tarr: Inorganic Chemistry, Pearson Publication.
10. Mahan, B.H. University Chemistry 3rd Ed. Narosa (1998).
11. Petrucci, R.H. General Chemistry 5th Ed. Macmillan Publishing Co., New York (1985).

Organic Chemistry

1. Organic Chemistry-P. Y. Bruice, 7th Edition, Pearson Education Pvt. Ltd., New Delhi (2013).

2. Heterocyclic Chemistry- R. K. Bansal, 3rd Edition, New- Age International, New Delhi,2004.
3. McMurry, J.E. Fundamentals of Organic Chemistry, 7thEd. Cengage Learning India Edition,2013.
4. Sykes, P. A Guidebook to Mechanism in Organic Chemistry, Orient Longman, New Delhi(1988).
5. Stereochemistry-Conformation and Mechanism-P. S. Kalsi, Wiley-Eastern Ltd, NewDelhi.
6. Morrison, R.T. & Boyd, R.N. Organic Chemistry, Pearson,2010.
7. Bahl, A. &Bahl, B.S. Advanced Organic Chemistry, S. Chand,2010.
8. Graham Solomons, T. W., Fryhle, C. B. & Snyder, S.A. Organic Chemistry, John Wiley & Sons (2014).
9. Organic Chemistry Volume-I, II- I. L. Finar, 6th Edition, ELBS London(2004).
10. Organic Chemistry-F. A. Carey, 4th Edition, McGraw Hill(2000).
11. Modern Organic Chemistry - R.O.C. Norman and D.J. Waddington, ELBS,1983.
12. Understanding Organic reaction mechanisms - A. Jacobs, Cambridge Univ. Press,1998.
13. Organic Chemistry - L. Ferguson, Von Nostrand,1985.
14. Organic Chemistry - M. K. Jain, Nagin & Co.,1987.
15. Organic Chemistry- Mehta and Mehta, PHI Learning Pvt. Ltd, NewDelhi,2005.

Physical Chemistry

1. Barrow, G.M. Physical Chemistry, Tata McGraw-Hill,2007.
2. Castellan, G.W. Physical Chemistry, 4th Ed. Narosa,2004.
3. Kotz, J.C., Treichel, P.M. & Townsend, J.R. General Chemistry Cengage Learning India Pvt. Ltd., New Delhi,2009.
4. P.W. Atkins: Physical Chemistry,2002.
5. W.J. Moore: Physical Chemistry,1972.
6. Text Book of Physical Chemistry - P. L. Soni, S. Chand & Co.,1993.
7. Text Book of Physical Chemistry - S. Glasstone, Mackmillan India Ltd.,1982.
8. Principles of Physical Chemistry - B. R. Puri, L. R. Sharma and M. S. Patania, S. L. N. Chand &Co. 1987.
9. Physical Chemistry - Alberty R. A. and Silbey, R. J. John Wiley and sons,1992.
10. Physical Chemistry - G. M. Barrow, McGraw Hill,1986.
11. Physical Chemistry (3rd Edition) - Gilbert W. Castilian, Narosa Publishing House,1985.
12. Chemical Kinetics by K. J. Laidler, Tata McGraw Hill Publishing Co., NewDelhi.
13. Kinetics and Reaction Mechanisms by Frost and Pearson, Wiley, New York,1981.

Open Elective Course - Semester – I

Title of the Course: OE-1: CHEMISTRY IN DAILY LIFE

Unit- I

Dairy Products: Composition of milk and milk products. Analysis of fat content, minerals in milk and butter. Estimation of added water in milk. Beverages: Analysis of caffeine in coffee and tea, detection of chicory in coffee, chloral hydrate in toddy, determination of methyl alcohol in alcoholic beverages. Food additives, adulterants, and contaminants- Food preservatives like benzoates, propionates, sorbates, disulphites. Artificial sweeteners: Aspartame, saccharin, dulcin, sucralose, and sodium cyclamate. Flavors: Vanillin, alkyl esters (fruit flavors), and monosodium glutamate. Artificial food colorants: Coal tar dyes and non-permitted colors and metallic salts. Analysis of pesticide residues in food. [14 hours]

Unit- II

Vitamins: Classification and Nomenclature. Sources, deficiency diseases, and structures of Vitamin A1, Vitamin B1, Vitamin C, Vitamin D, Vitamin E & Vitamin K1. Oils and fats: Composition of edible oils, detection of purity, rancidity of fats and oil. Tests for adulterants like argemone oil and mineral oils. Halphen test. Soaps & Detergents: Definition, classification, manufacturing of soaps and detergents, composition and uses [14 hours]

Unit- III

Chemical and Renewable Energy Sources: principles and applications of primary & secondary batteries and fuel cells. Basics of solar energy, future energy storer. Polymers: Basic concept of polymers, classification and characteristics of polymers. Applications of polymers as plastics in electronic, automobile components, medical fields, and aerospace materials. Problems of plastic waste management. Strategies for the development of environment-friendly polymers. [14hours]

COURSE OUTCOMES OEC-1 Chemistry

On completion of the course students will be able to:

1. Understand the chemical constituents in various day to day materials using by a commonman.
2. Understand the chemical constituents in vitamins, soaps and detergents
3. Understand the renewable chemical energy resources
4. Understand different types of polymers and their applications.

Reference Books

1. B. K. Sharma: Introduction to Industrial Chemistry, Goel Publishing, Meerut (1998)
2. Medicinal Chemistry- Ashtoush Kar.
3. Analysis of Foods – H.E. Cox: 13.
4. Chemical Analysis of Foods – H.E. Cox and Pearson.
5. Foods: Facts and Principles. N. Shakuntala Many and S. Swamy, 4th ed. New Age International (1998)
6. Physical Chemistry – P I Atkins and J. de Paula – 7th Ed. 2002, Oxford University Press.
7. Handbook on Fertilizer Technology by Swaminathan and Goswamy, 6th ed. 2001, FAI.
8. Organic Chemistry by I. L. Finar, Vol. 1 & 2. 9. Polymer Science and Technology, J. R. Fried (Prentice Hall).

Open Elective Course - Semester – II

Title of the Course: OE-2: Molecules of Life

UNIT I

Carbohydrates

Classification of carbohydrates, reducing and non-reducing sugars, General properties of glucose and fructose, their open chain structures. Epimers, mutarotation and anomers.

Linkage between monosaccharides, structure of disaccharides (sucrose, maltose, lactose) and polysaccharides (starch and cellulose) excluding their structure elucidation.

Amino Acids, Peptides and Proteins

Classification of amino acids, Zwitterion structure and Isoelectric point. Overview of Primary, Secondary, Tertiary and Quaternary structure of proteins. Determination of primary structure of peptides.

UNIT II

Enzymes and correlation with drug action

Mechanism of enzyme action, factors affecting enzyme action, Co-enzymes and cofactors and their role in biological reactions, Specificity of enzyme action (including stereospecificity),

Enzyme inhibitors and their importance, phenomenon of inhibition (Competitive and Non competitive inhibition including allosteric inhibition).

Drug action-receptor theory. Structure–activity relationships of drug molecules, binding role of –OH group, –NH₂ group, double bond and aromatic ring

Lipids

Introduction to lipids, classification. Biological importance of triglycerides, phospholipids, glycolipids, and steroids (cholesterol).

UNIT III

Nucleic Acids

Components of nucleic acids: Adenine, guanine, thymine and cytosine (Structure only), other components of nucleic acids, Nucleosides and nucleotides (nomenclature), Structure of polynucleotides; Structure of DNA (Watson-Crick model) and RNA (types of RNA), Genetic Code, Biological roles of DNA and RNA: Replication, Transcription and Translation.

Concept of Energy in Biosystems

Calorific value of food. Standard caloric content of carbohydrates, proteins and fats. Oxidation of foodstuff (organic molecules) as a source of energy for cells. Introduction to Metabolism (catabolism, anabolism), ATP: the universal currency of cellular energy, ATP hydrolysis and free energy change. Conversion of food into energy. Outline of catabolic pathways of Carbohydrate- Glycolysis, Fermentation, Krebs Cycle. Overview of catabolic pathways of Fats and Proteins. Interrelationships in the metabolic pathways of Proteins, Fats and Carbohydrates

Course Outcome / Learning Outcome:

After studying this paper the student would be able to

1. Acquire knowledge about different types of sugars and their chemical structures.
2. Identify different types of amino acids and determine the structure of peptides.
3. Explain the actions of enzymes in our body and interpret enzyme inhibition.
4. Predict action of drugs. Depict the biological importance of oils and fats. Importance of lipids in the metabolism. Differentiate RNA and DNA and their replication. Explain production of energy in our body.

Reference Books:

1. Morrison, R. T. & Boyd, R. N. Organic Chemistry, Dorling Kindersley (India) Pvt.Ltd. (Pearson Education).
2. Finar, I. L. Organic Chemistry (Volume 1), Dorling Kindersley (India) Pvt. Ltd.(Pearson Education).
3. Finar, I. L. Organic Chemistry (Volume 2), Dorling Kindersley (India) Pvt. Ltd.(Pearson Education).
4. Nelson, D. L. & Cox, M. M. Lehninger's Principles of Biochemistry 7th Ed.,
5. W. H. Freeman. Berg, J.M., Tymoczko, J.L. & Stryer, L. Biochemistry, 2002.